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292 K? $n = PV = (2.8 \text{ atm})(98 \text{ L}) = 11$
moles of gas $RT (0.0821$

$\text{L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})$ 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25 0

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Solutions 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K? $n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821$

$\text{L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})$ 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25 0

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How many moles of gas occupy 98 L

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at a pressure of 2.8 atmospheres and a temperature of 292 K? $n = PV = (2.8 \text{ atm})(98 \text{ L}) = 11 \text{ moles of gas}$ $RT (0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})$ 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are

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Gas Laws Worksheet - Solutions 1)

How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K? $n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11 \text{ moles of gas}$

2) If 5.0

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