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$P_{\text{vap}} = \exp\left(\frac{14.790}{T} - K\right)$  and by the ideal gas law  
 $f_{\text{Solutions to Chemical and Engineering Thermodynamics,}}$   
 $3e N V R T \times 18 \text{ mol} \times 8.314 \times 10^{-5} (\text{bar} \quad \text{m}^3 \text{ mol}^{-1} \text{ K}^{-1}) T P =$   
 $= V V 1 \times 10^{-3} - (10 - x) 106$  volume taken up by liquid  
Equating  $P$  and  $P_{\text{vap}}$  we have  $x 18 \times 8.314 \times 10^{-5} T F$   
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