

Chemical Biochemical And Engineering Thermodynamics Sandler Solution Manual

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$$P_{\text{vap}} = \exp\left(\frac{14.790}{T} - \frac{5713}{T^2}\right)$$
 and by the ideal gas law $f = \frac{P_{\text{vap}}}{P} = \frac{V_{\text{vap}}}{V}$
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$$P_{\text{vap}} = \exp\left(\frac{14.790}{T} - \frac{5713}{T^2}\right)$$

$$P_{\text{vap}} = \exp\left(\frac{14.790}{106} - \frac{5713}{106^2}\right)$$

106 volume taken up by liquid Equating P and P_{vap} we have
$$x = \frac{P_{\text{vap}}}{P} = \frac{5432.8 \text{ bar}}{18 \times 8.314 \times 10^5 \text{ (bar} \cdot \text{m}^3 \text{ mol}^{-1} \text{K}^{-1})}$$

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