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- No two electrons in an atom can have the same unique set of four quantum numbers. •
- The first quantum number describes the principal energy level. For example, the quantum number 2 identifies the second principal energy level. •

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The first two quantum numbers together describe a sublevel.

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List the three statements that make up the modern atomic theory. 2. Define atomic number. What is the atomic number for a boron atom? 3. Which elements are alkali metals, alkaline earth metals, and halogens? 4. Two of the most well-known isotopes of carbon are carbon-12 and carbon-14.

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Objective 3.11.1 The  
Mysterious Electron 417 the  
electron are also called  
orbitals. The orbital shown  
in Figure 11.3 is called the  
1s orbital. The negative  
charge is most intense at  
the nucleus and decreases in  
intensity with distance  
outward.

### **Chapter 11 Modern atomic theory**

Chapter 4 Atomic Structure  
Section 4.3 Modern Atomic  
Theory (pages 113-118). Use  
this table to answer  
questions 12 and 13. Name  
Chapter 4: Atomic Structure  
Worksheet. Answer the  
following questions, circle  
the best answer.

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The modern atomic theory states that atoms of one element are the same, while atoms of different elements are different. What makes atoms of different elements different? The fundamental characteristic that all atoms of the same element share is the number of protons. All atoms of hydrogen have one and only one proton in the nucleus; all atoms of iron have 26 protons in the nucleus.

**Atomic Theory - Introductory  
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Hydrogen Atoms Particle  
Interpretation of the Wave  
Character of the Electron  
Other Important Waveforms  
Overall Organization of  
Principal

### **Chapter 11 Modern Atomic Theory**

separate atoms (Figure 4.1).  
Radiant energy from the sun  
causes sunburn 4.1 Energy  
Some reactions absorb  
energy. Energy from the sun  
 $O_2(g) + \text{energy} \rightarrow 2O(g)$  Figure  
4.1 Some reactions absorb



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energy. 120 Chapter 4 Modern Atomic Theory Before we can begin to explain the role that energy plays in these and other chemical

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culmination of the authors' many years of teaching and research in atomic physics, nuclear and particle physics, and modern physics. It is also a crystallization of their intense passion and strong interest in the history of physics and the philosophy of science. Together with the solution manual which presents solutions to many end-of-chapter problems in the textbook, they are a valuable resource to the instructors and students working in the modern atomic field."--Publisher's website.

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Alchemists are generally held to be the quirky forefathers of science, blending occultism with metaphysical pursuits. Although many were intelligent and well-intentioned thinkers, the oft-cited goals of alchemy paint these antiquated experiments as wizardry, not scientific investigation. Whether seeking to produce a miraculous panacea or struggling to transmute lead into gold, the alchemists radical goals held little relevance to consequent scientific pursuits. Thus, the temptation is to view the transition from alchemy

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to modern science as one that discarded fantastic ideas about philosophers stones and magic potions in exchange for modest yet steady results. It has been less noted, however, that the birth of atomic science actually coincided with an efflorescence of occultism and esoteric religion that attached deep significance to questions about the nature of matter and energy. Mark Morrisson challenges the widespread dismissal of alchemy as a largely insignificant historical footnote to science by prying into the revival of alchemy and its influence on the emerging subatomic

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sciences of the late 19th and early 20th centuries. Morrison demonstrates its surprising influence on the emerging subatomic sciences of the late 19th and early 20th centuries. Specifically, Morrison examines the resurfacing of occult circles during this time period and how their interest in alchemical tropes had a substantial and traceable impact upon the science of the day. Modern Alchemy chronicles several encounters between occult conceptions of alchemy and the new science, describing how academic chemists, inspired by the alchemy



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Mark Bishop attempted to transmute the elements; to make gold. Examining scientists publications, correspondence, talks, and laboratory notebooks as well as the writings of occultists, alchemical tomes, and science-fiction stories, he argues that during the birth of modern nuclear physics, the trajectories of science and occultism---so often considered antithetical---briefly merged.

Each text in this series provides a concise account of the basic principles underlying a given subject,

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embodying an independent-learning philosophy and including worked examples. This text covers atomic structure and periodicity.

Steve and Susan Zumdahl's texts focus on helping students build critical thinking skills through the process of becoming independent problem-solvers. They help students learn to think like a chemists so they can apply the problem solving process to all aspects of their lives. In CHEMISTRY: AN ATOMS FIRST APPROACH, the Zumdahls use a meaningful approach that

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begins with the atom and proceeds through the concept of molecules, structure, and bonding, to more complex materials and their properties. Because this approach differs from what most students have experienced in high school courses, it encourages them to focus on conceptual learning early in the course, rather than relying on memorization and a plug and chug method of problem solving that even the best students can fall back on when confronted with familiar material. The atoms first organization provides an opportunity for students to use the tools of critical

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thinkers: to ask questions, to apply rules and models and to evaluate outcomes.

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